

QP CODE:101006

Reg.No:

First Year B.Pharm Degree Supplementary Examinations - March 2014

(2010 - Scheme)

PHARMACEUTICAL CHEMISTRY - I
(Inorganic & Physical Chemistry)

Time: 3 Hours

Total Marks: 100

- Answer all Questions.
- Write equations wherever necessary.

Essay

(3x10=30)

1. Mention the importance of physiological acid base balance with a detailed note on the electrolytes used in acid base therapy.
2. Define adsorption. Describe the adsorption isotherms and various factors influencing adsorption.
3. List the importance of limit test and explain how it is performed in case of iron and arsenic.

Short notes

(14x5=70)

4. What are colligative properties. Explain Debye-Huckel theory.
5. Discuss on surface tension and what is its importance in Pharmacy.
6. Explain briefly the following co-ordination compounds
 - Dimercaprol
 - Penicillamine
7. Mention the medicinal gases commonly used. How nitrous oxide is assayed.
8. Define the term astringent. Add a note on alum and zinc chloride.
9. Classify antacids and add a note on milk of magnesia.
10. Discuss on any two radio pharmaceutical preparations with its clinical application.
11. Define refractive index. Explain the working of Abbe refractometer.
12. What is the role of fluorides as anti-carries agent. Explain with suitable examples.
13. Explain the preparation, assay and uses of hydrogen peroxide.
14. Describe the assay of calcium gluconate.
15. Explain optical activity. Describe the working of polarimeter.
16. Sclerosing agents.
17. Complete the following equation:
 - $2\text{KMnO}_4 + 3\text{H}_2\text{SO}_4 =$
 - $\text{Na}_2\text{B}_4\text{O}_7 + \text{H}_2\text{SO}_4 + 5\text{H}_2\text{O} =$
 - $\text{NaOH} + \text{H}_2\text{CO}_3 =$
 - $\text{NH}_4\text{Cl} + \text{HCHO} =$
 - $\text{NaF} + \text{HCl} =$
